

| Physical Science Essentials I | Scope and Sequence |
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| Unit Lesson | Objectives |
| Energy and Motion | |
| Speed and Velocity | |
| | Differentiate between speed and velocity. |
| | Solve problems involving distance, time, speed, and/or velocity. |
| | Interpret graphs of distance versus time. |
| Acceleration | |
| | Describe the concept of acceleration. |
| | Solve problems involving velocity, time, and acceleration. |
| | Interpret graphs of velocity versus time. |
| Lab: Motion | |
| | Measure distance and time to determine speed. |
| | Graph changes in motion. |
| | Interpret data to determine acceleration. |
| Potential and Kinetic Energy | |
| | Distinguish between potential and kinetic energy. |
| | Calculate the potential energy in a system. |
| | Calculate the kinetic energy in a system. |
| | Explain how energy is transferred in a moving system. |
| Lab: Kinetic Energy | |

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| Unit | Lesson | Objectives |
| | | Calculate the kinetic energy of objects of different mass. |
| | | Determine the kinetic energy of objects at different speeds. |
| | | Graph data to illustrate changes in kinetic energy. |
| | Unit Test | |
| Wave | 2S | |
| | Sound Waves | |
| | | Describe how sound waves are produced and how they travel. |
| | | Identify the features of a sound wave. |
| | | Explain how different materials and different temperatures affect the speed of sound waves. |
| | Hearing and the Ear | |
| | | Identify the parts of the ear. |
| | | Describe the function of each section of the ear. |
| | | Examine the causes of hearing loss. |
| | Using Sound | |
| | | Explain how and why animals use echolocation. |
| | | Describe the uses of ultrasound technology. |
| | | Summarize the ways in which sound waves are used for communication. |
| | The Electromagnetic Spectrum | |
| | | Describe the different parts of the electromagnetic spectrum. |
| | | Distinguish how electromagnetic waves differ from one another. |

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| Unit | Lesson | Objectives |
| | | Identify how different types of electromagnetic waves are used. |
| | Reflection and Mirrors | |
| | | Explain how light is reflected from a surface. |
| | | Describe the law of reflection. |
| | | Describe how a mirror forms an image. |
| | | Identify the types of images formed by different kinds of mirrors. |
| | Unit Test | |
| Elements and the Periodic Table | | |
| | Metals | |
| | | Describe the characteristic properties of metals. |
| | | Identify the location of metals in the periodic table. |
| | | Explain how and why the reactivity of metals changes in the periodic table. |
| | Nonmetals | |
| | | Describe the characteristic properties of nonmetals. |
| | | Identify the location of nonmetals in the periodic table. |
| | | Explain how and why the reactivity of nonmetals changes in the periodic table. |
| | Metalloids | |
| | | Describe the characteristic properties of metalloids. |
| | | Identify the location of metalloids in the periodic table. |
| | | Explain why most metalloids are used as semiconductors. |
| | Radioactivity | |
| | | Summarize the discovery of radioactivity. |

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| | | Explain radioactive decay. |
| | | Perform calculations involving the half-life of a radioactive isotope. |
| | | Describe applications of radioactive isotopes. |
| | Chemical Bonding | |
| | | Explain why atoms bond. |
| | | Identify the three types of bonds. |
| | | Complete electron dot diagrams. |
| | Unit Test | |
| Physical and Chemical Reactions | | S |
| | Balancing Chemical Equations | |
| | | Explain what it means for a chemical equation to be balanced. |
| | | Demonstrate how to balance a chemical equation. |
| | | Relate balanced chemical equations to the law of conservation of mass. |
| | Rate of Chemical Reactions | |
| | | Explain activation energy and its importance to chemical reactions. |
| | | Describe the factors that affect the rate of a chemical reaction. |
| | | Recognize how a catalyst and an inhibitor affect a chemical reaction. |
| | Lab: Rate of Chemical | |

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| | Reactions | |
| | | Describe the signs of a chemical reaction. |
| | | Identify how temperature and surface area affect the rate of a chemical reaction. |
| | | Science Practice: Conduct several controlled tests of multiple variables using repeated trials during an investigation about chemical reaction rate. |
| | Separation of Mixtures | |
| | | Explain why mixtures are able to be separated. |
| | | Describe methods for separating mixtures. |
| | Properties of Acids and Bases | |
| | | Compare the properties of acids and bases. |
| | | Describe common uses for acids and bases. |
| | Unit Test | |
| Properties of Matter | | |
| | The Four States of Matter | |
| | | Describe the arrangement and motion of atoms in the different states of matter. |
| | | Discriminate the characteristics of solids, liquids, gases, and plasma. |
| | Chemical Changes | |
| | | Describe and give examples of chemical properties of matter. |
| | | Explain what happens during a chemical change. |
| | | Describe the evidence that shows a chemical change has occurred. |
| | | Compare and contrast the properties of a new substance with the original substance after a chemical change. |
| | Density | |

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| | Explain density and state the SI units used to measure it. |
| | Calculate the mass, volume, or density of an object given the other two measurements. |
| | Determine whether an object will sink or float relative to the density of the surrounding liquid. |
| Lab: Density of Solids | |
| | Measure the mass and volume of various solid objects. |
| | Calculate the density of several solid objects. |
| | Use density to identify an unknown substance. |
| Behavior of Gases | |
| | Distinguish between Boyle's law, Charles's law, and Gay-Lussac's law. |
| | Use the combined gas law to determine pressure, temperature, and volume of a gas. |
| Unit Test | |
| Cumulative Exam | |
| Cumulative Exam Review | |
| Cumulative Exam | |